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Chapter 11 187 Exercise 11.3 □

Equation Stoichiometry: Iron is combined with carbon in a series of reactions to form pig iron, which is about 4.3% carbon.

$$2C + O_2 \rightarrow 2CO$$
$$Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$$

2CO + C (in

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iron)  $\text{CO}_2$  Pig iron is easier to shape than pure iron, and the presence of carbon lowers its melting point

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AP Chemistry □ Solutions, Chapter 11/4b FAQ. Q: What are hydrophobic

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and hydrophilic? A: Hydrophobic is a "fear of water". Nonpolar substances are typically hydrophobic. They are not attracted to water. Hydrophilic on the other hand are "water lovers". These compounds are attracted to water, and therefore are commonly ionic or polar.

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Strings Electrons as Standing Waves

Waveforms for Hydrogen Atoms

Particle Interpretation of the Wave

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Character of the Electron Other Key

Important Waveforms Overall

Organization of Principal Energy

Levels, Sublevels, and Orbitals 11.2

Multi-Electron Atoms.

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PLAY. Match. ... Terms in this set (33)

Stoichiometry. study of quantitative relationships between the amounts of reactants used and the amounts of products formed in a chemical reaction

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Stoichiometry Unit 4 Study Guide. 23  
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Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon

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dioxide and water 12. 160.10 g 13.  
160.10 g 14. They are equal. 15. A  
mole ratio is a ratio between ...

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chapter 13 (chemical kinetics), chapter 14 (chemical equilibrium), and chapter 21 (nuclear chemistry). all of this content will be on the 2nd exam on march 11

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Exercise 11.3 □ Equation

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from the primary book and enhance

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important to understanding the processes in an industrial plant. For biologists and health professionals, it is essential because nearly all of biochemistry springs from organic chemistry. Additionally, all scientists can benefit from improved critical thinking and problem-solving skills that

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are developed from the study of organic chemistry. Organic chemistry, like any "skill", is best learned by doing. It is difficult to learn by rote memorization, and true understanding comes only from concentrated reading, and working as many problems as possible. In fact, problem



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adjacent chapters. This approach helps students see the strong connections that exist between these three branches of chemistry, and allows instructors to discuss these, interrelationships while the material is still fresh in students' minds. This integration involves the following sets

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of chapters: \* Chapter 3 (Compounds) and Chapter 4 (An Introduction to Organic Compounds). An introduction to bonding and compounds is followed by a look at the members of a few key organic families. \* Chapters 3, 4 and 6.(Reactions).. A study of inorganic.and organic compounds is

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followed (after a look at gases, liquids, and solids in Chapter 5) by an introduction to their reactions. \*

Chapter 7 (Solutions) and Chapter 8 (Lipids and Membranes) A discussion of solubility is followed by a look at the importance of solubility in biochemistry. Some reactions from

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Chapter 6 are reintroduced. \* Chapter 9.(Acids and Bases) and Chapter 10 (Carboxylic Acids, Phenols and Amines) Principles of acid/base Chemistry from an inorganic perspective are followed by a chapter on the organic and biochemical aspects of this topic. \* Chapter 11

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(Alcohols, Aldehydes and Ketones) and Chapter 12 (Carbohydrates). An introduction to the chemistry of alcohols, aldehydes and ketones is followed by a presentation of related biochemical applications.

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